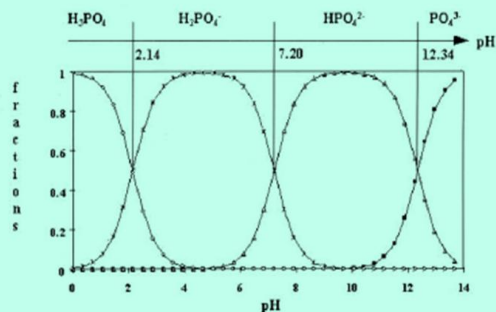


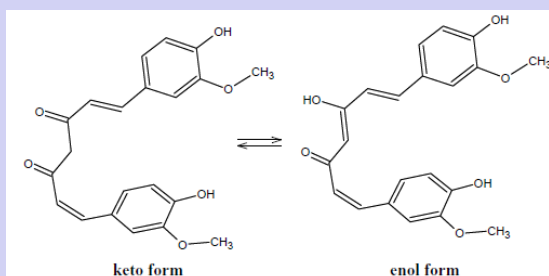
**Table 2. Dismutation Equilibria of the Ampholytes in the Systems**

Proton-Phosphates, Manganese(II)-Oxalates (Ox) and Iron(II)-*ortho*-Phenanthroline (*o*-phen) with Corresponding Values of the log of the Equilibrium Constants

Ampholyte	Dismutation Equilibrium	log K
HPO <sub>4</sub> <sup>2-</sup>	2HPO <sub>4</sub> <sup>2-</sup> = H <sub>2</sub> PO <sub>4</sub> <sup>-</sup> + PO <sub>4</sub> <sup>3-</sup>	-5.1
	3HPO <sub>4</sub> <sup>2-</sup> = H <sub>3</sub> PO <sub>4</sub> + 2PO <sub>4</sub> <sup>3-</sup>	-15.3
H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>	2H <sub>2</sub> PO <sub>4</sub> <sup>-</sup> = H <sub>3</sub> PO <sub>4</sub> + HPO <sub>4</sub> <sup>2-</sup>	-5.1
	3H <sub>2</sub> PO <sub>4</sub> <sup>-</sup> = 2H <sub>3</sub> PO <sub>4</sub> + PO <sub>4</sub> <sup>3-</sup>	-15.3
MnOx	2MnOx = MnOx <sub>2</sub> <sup>-</sup> + Mn <sup>2+</sup>	0.8
	3MnOx = MnOx <sub>3</sub> <sup>4-</sup> + 2Mn <sup>2+</sup>	-1.2
MnOx <sub>2</sub> <sup>2-</sup>	2MnOx <sub>2</sub> <sup>2-</sup> = MnOx <sub>3</sub> <sup>4-</sup> + MnOx	-2.9
	3MnOx <sub>2</sub> <sup>2-</sup> = 2MnOx <sub>3</sub> <sup>4-</sup> + Mn <sup>2+</sup>	-5.0
Fe( <i>o</i> -phen) <sup>2+</sup>	2Fe( <i>o</i> -phen) <sup>2+</sup> = Fe( <i>o</i> -phen) <sub>2</sub> <sup>2+</sup> + Fe <sup>2+</sup>	-0.5
	Fe( <i>o</i> -phen) <sup>2+</sup> = Fe( <i>o</i> -phen) <sub>2</sub> <sup>2+</sup> + 2Fe <sup>2+</sup>	3.78
	2Fe( <i>o</i> -phen) <sub>2</sub> <sup>2+</sup> = Fe( <i>o</i> -phen) <sub>3</sub> <sup>2+</sup> + Fe( <i>o</i> -phen) <sup>2+</sup>	4.7
Fe( <i>o</i> -phen) <sub>2</sub> <sup>2+</sup>	3Fe( <i>o</i> -phen) <sub>2</sub> <sup>2+</sup> = Fe( <i>o</i> -phen) <sub>3</sub> <sup>2+</sup> + Fe <sup>2+</sup>	9.0



J. Chem. Educ. 1995, 72, 1099-1105.

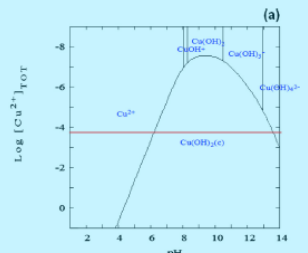
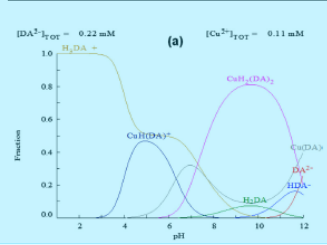


Spectrochim. Acta A. 2004, 60, 1091-1097.

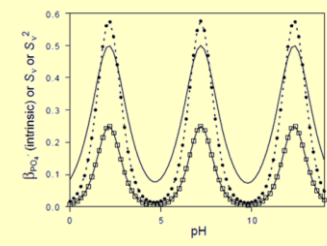
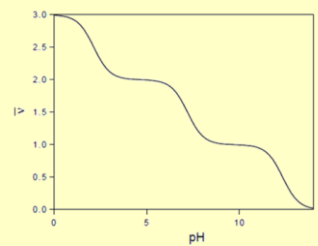
**Table 3**  
Comparison between pK<sub>A</sub> values reported by Jovanovic et al. [15], and those obtained in the present work using SQUAD

Equilibrium <sup>a</sup>	[15]	SQUAD
HCur = H <sup>+</sup> + Cur	10.41 ± 0.05	10.51 ± 0.01
H <sub>2</sub> Cur = H <sup>+</sup> + HCur	8.55 ± 0.05	9.88 ± 0.02
H <sub>3</sub> Cur = H <sup>+</sup> + H <sub>2</sub> Cur		8.38 ± 0.04

**Table 1**  
Global formation constants for the DA/Cu(II) species obtained through point analysis using SQUAD, with a standard deviation (σ) of 1.44 × 10<sup>-2</sup> and a quadratic sum of the absorbance residues (U) of 4.08 × 10<sup>-1</sup>.



Spectrochim. Acta A. 2015, 143, 187-191.



J. Chem. Educ. 2002, 79, 389-392.

**Table 1**  
Global stability constants of dimeric and monomeric Cu(II)/L species. σ<sub>ij</sub> is the standard deviation of the constant, U is the sum of squared residuals.

L	Stoichiometry	[Cu <sub>2</sub> L <sub>2</sub> ] <sup>2+</sup>	logβ <sub>ij</sub>	σ <sub>ij</sub>	U
AcO	2:2	[Cu <sub>2</sub> AcO <sub>2</sub> ] <sup>2+</sup>	12.81	0.22	3.25 × 10 <sup>-1</sup>
	2:3	[Cu <sub>2</sub> AcO <sub>3</sub> ] <sup>+</sup>	17.29	0.28	
	2:4	Cu <sub>2</sub> AcO <sub>4</sub>	21.08	0.28	
Dic	2:2	[Cu <sub>2</sub> Dic <sub>2</sub> ] <sup>2+</sup>	12.96	0.29	4.55 × 10 <sup>-1</sup>
	2:3	[Cu <sub>2</sub> Dic <sub>3</sub> ] <sup>+</sup>	17.14	0.49	
	2:4	Cu <sub>2</sub> Dic <sub>4</sub>	21.47	0.41	
AcO	1:2	CuAcO <sub>2</sub>	8.49	0.02	2.01 × 10 <sup>-1</sup>
Dic	1:2	CuDic <sub>2</sub>	9.15	0.04	7.67 × 10 <sup>-2</sup>

Polyhedron. 2021, 209, 115486.

